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Question 4.4
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Question 4.6
Question 4.7
Question 4.8
Question 4.9
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Question 4.40

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NCERT Q 4.1

NCERT Q 4.2

NCERT Q 4.3

NCERT Q 4.4

NCERT Q 4.5

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NCERT Q 4.7

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Introduction

Exercise 4.1

Exercise 4.2

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Exercise 4.4

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Question 7.30

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Atoms

Kossel-Lewis approach

Lewis Symbols

Lewis Dot Structures

Lewis Dot Structure: H₂O

Lewis Dot Structure:C2H2

Lewis Dot Structure:Nitrite ion

Lewis Dot Structure: HNO₃

Example on Formal Charge

Limitation of Octet Rule

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Introduction: Structure of Atom

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Introduction

Atom

Molecules Formation Theories

Kossel-Lewis Approach

Lewis Symbols

Octet Rule

Lewis Dot structure:CO₂

Charge on Ions

F.C :Ex 1

Why F.C

Limitation:Octet Rule

Ionic Bond

I.E.

Electron Gain Enthalpy

Lattice Enthalpy

Ex:NaCl

Stable Ionic Bond

Bond Length

Atomic radius

Vander Waals Radii

Bond Enthalpy

Bond Angle

Bond Order

Resonance Struc.

Bond Polarity

D.M:PolyAtomic

Fajan's Rule

VSEPR Theory

Postulate 1

Postulate 2

Postulate 3

Postulate 4

Geometry VSEPR

Shape vs.Geometry

VSEPR theory:CO₂

VSEPR:Merits

VSEPR:Limitations

VBT

Orbital Overlap

Orbital Overlap Types

Covalent Bond types

Sigma Bond

Pi bonds

Str.:Å & Å

Rotation

Å/s. Å

Draw Shapes

Hybridisation

Hybridisation:sp

Hybridisation:sp²

Hybridisation:sp³

Hybridisation:d orb.

VBT:Limitations

MOT

MOT formation

LCAO conditions

MO:Types

Energy Dia. :MO
Molecular Behaviour
Bond Order
StabilityMolecules
Bond Length
Magnetic nature
Hydrogen Bonding
Covalent Vs.H bond

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Ques 11 to 17

Ques 18 to 23

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32 minutes - VSPER Theory + Molecular Orbital Theory [MOT] + Hybridization Theory + Valence
Bond Theory - Everything in a Single Video in ...

intro

Octet Rule

Types of Bond

Octet Rule Limitation

Bond Pair & Lone Pair

Method 1 (Find Structures)

Method 2 (Formula Method)

Most Stable Structure

Valence Bond Theory (VBT)

Limitation in VBT

Molecular Orbital Theory

Paramagnetic vs Diamagnetic

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