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General Chemistry II - Aqueous Ionic Equilibrium - Ch 18a - General Chemistry II - Aqueous Ionic Equilibrium - Ch 18a by Chemistry Student Resource Channel 2,063 views 2 years ago 55 minutes - All right so welcome to **chapter 18**, which will be aqueous ionic equilibrium you could also refer to this chapter as buffers um but ...

Chapter 18 Solution Part 3/7 - Chapter 18 Solution Part 3/7 by CK 351 views 3 years ago 15 minutes - Calculation of solubility Example 1 At 20°C, 50g of potassium nitrate saturated **solution**, is evaporated to get 12g of potassium ...

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Part B: Calculate the pH at the equivalence point
Part C: Calculate the pH at 50 mL

Chapter 18-2 Buffers - Chapter 18-2 Buffers by Louise Thompson 116 views 10 years ago 14 minutes, 14 seconds - Chapter 18,-2 Buffers.

A buffer causes solutions to resist a change in pH when a strong acid or base is added Add strong acid or base to blood (-0.01 mol)

Two requirements for a buffer: Two substances are needed: an acid that is capable of reacting with added OH ions and a base that can consume

How do buffers work? Consider acetic acid/acetate ion

pH of a buffer solution is controlled by two factors The strength of the acid (expressed by both K, or pk.) The relative amounts of acid and conjugate base

pH control If you rearrange the K expression, you get the following equation which shows mathematically how the pH can be adjusted

Buffer capacity The buffer should have the capacity to control the pH after the addition of reasonable amounts of acid and base. Acid concentration must be enough to consume all added OH and still control pH

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