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divide each number by the smallest of these three values
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take the molar mass of the molecular formula and divide
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Electrolysis
Exothermic and Endothermic
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Balancing a Combustion Reaction
Limiting Reactant
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Convert Moles into Grams
Percent Yield
Find the Percent Error
Percent Error Equation
The Amount of Excess Reactant That Remains
Limiting Reactant and Convert It to the Grams of the Excess Reactant
Molar Ratio
Convert Moles of C_2H_6 into Grams
Identify the Limiting Reactant
The Theoretical Yield
Convert Moles of Ethanol into Moles of the Product CO_2
Stoichiometric Relationship between the Grams of Oxygen Gas and Carbon Dioxide
Calculate the Actual Yield
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finding the empirical formula from the mass of CO_2
find the empirical formula of C_4H_8
start with 20 grams of carbon
divide each number by the lowest number
calculate the molar mass of the empirical formula
find the empirical formula
convert the grams of every element

know the molar mass of carbon
need to multiply the subscripts by a whole number
multiply the subscripts by 3
find the molar mass of the empirical form
find the molecular formula
find the empirical formula of the compound
find the number of moles of carbon
start with the grams of CO_2
find the moles of carbon
molecular formula has a molar mass of 216
find the molar mass of the empirical
take the molar mass of the molecular formula
determine the empirical form of the compound
find the moles of oxygen from CO_2 and water
find the moles of carbon and hydrogen
start with the eight point nine five two grams of CO_2
get the grams of oxygen
start with the point two zero three five moles of carbon
find the mass of oxygen
convert grams of oxygen into moles
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Polyphenyl Ethene
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start
Atomic Structure, Ar and the Periodic Table
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History of the Atom
Groups, Periods and Electronic Structure
Ions
The Early Periodic Table, Newlands and Mendeleev
Group 0/Group 8/Noble Gases
Group 7/The Halogens
Group 1/The Alkali Metals
States of Matter, Boiling Points, Melting Points and State Symbols
Ionic Bonding
Covalent Bonding and Dot and Cross Diagrams
Structural Formula
Metals and Alloys
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Ionic Compound Structure and Properties
Simple Covalent Compound Structure and Properties
Giant Covalent Structure and Properties
Diamond, Graphite and their Comparisons
Fullerenes and the Comparison of Carbon Nanotubes and Buckminsterfullerene
Polymers

Measuring Reactants and Products and Degrees of Uncertainty
Concentration equation
Building Symbol Equations
Working Out Mr
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Reacting Masses
Limiting Reactions and Limiting Reactants
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Making a Pure Salt Method/Experiment
pH Scale and Neutralisation
Titration Method/Experiment
Strong Acids, Weak Acids and the Concentration of Acids
Conductivity of Ionic Compounds
Electrolysis of Aluminium
Common Setups for Electrolysis and Testing for Hydrogen, Oxygen and Chlorine
Exothermic and Endothermic Reactions
Bond Energy
End of 9-1 Combined Science Content
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Hydrogen Fuel Cells
Transition Metals
Nanotechnology/Nanoscience
Percentage Yield
Atom Economy
Titration Calculations
Mole of a Gas
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Titration
Indicators
Sodium Chloride
Electrolysis
Periodic Table
Balanced Symbol Equation
Distillation
Evaporation
Periods
Electronic Configuration
Sodium

Ionic Bonding
Magnesium Oxide
Covalent Bonding
Formula of the Salts
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Going Over The Mark Scheme and Examiners Report
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Acidic Buffer
Basic Buffers
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Chromatography
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Question Eight
Magnesium Nitride
Question 10 the Electrolysis of Constrained Hydrochloric Acid
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